

The Mole

What is the difference between "molar mass" and "atomic mass"?	Molar Mass is used to describe the mass of one mole of a chemical compound. Atomic mass is used to describe the mass of one mole of an element or the mass of one atom of an element.
How many moles are present in 27 grams of $\text{Cu}(\text{OH})_2$?	.28 mol
How many moles are present in 2.45×10^{24} molecules of CH_4 ?	4.07 mol
How many grams are there in 3.4×10^{24} molecules of NH_3 ?	84.8 g
How much does 2.4 moles of $\text{Ca}(\text{NO}_3)_2$ weigh?	393.83 g
What is the molar mass of MgO ?	40.3 g/mol
How many molecules are there in 41 grams of FeF_3 ?	2.19×10^{23} molecules
How many molecules are there in 225 grams of Na_2SO_4 ?	9.53×10^{23} molecules
How many grams are there in 2.3×10^{23} atoms of silver?	41.2 g
How many grams are there in 7.4×10^{23} molecules of AgNO_3 ?	208.81 g